

**Stoichiometry limiting reagent worksheet answer key**

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Take the reaction:  $\text{NH}_3 + \text{O}_2 \rightarrow \text{NO} + \text{H}_2\text{O}$ . In an experiment, 3.25 g of  $\text{NH}_3$  are allowed to react with 3.50 g of  $\text{O}_2$ .  
a. Which reactant is the limiting reagent?  
b. How many grams of  $\text{CO}_2$  are formed?  
c. How much of the excess reactant remains after the reaction?

Take the reaction:  $\text{C}_6\text{H}_6 + \text{Br}_2 \rightarrow \text{C}_6\text{H}_5\text{Br} + \text{HBr}$ .  
a. What is the theoretical yield of  $\text{C}_6\text{H}_5\text{Br}$  if 42.1 g of  $\text{C}_6\text{H}_6$  reacts with 73.0 g of  $\text{Br}_2$ ?  
b. If the actual yield of  $\text{C}_6\text{H}_5\text{Br}$  is 63.6 g, what is the percent yield? 88.8%  
Use the following reaction:  
 $\text{C}_4\text{H}_9\text{OH} + \text{NaBr} + \text{H}_2\text{SO}_4 \rightarrow \text{C}_4\text{H}_9\text{Br} + \text{NaHSO}_4 + \text{H}_2\text{O}$   
If 15.0 g of  $\text{C}_4\text{H}_9\text{OH}$  react with 22.4 g of  $\text{NaBr}$  and 32.7 g of  $\text{H}_2\text{SO}_4$  to yield 17.1 g of  $\text{C}_4\text{H}_9\text{Br}$ , what is the percent yield of this reaction? 61.6%  
Silicon nitride ( $\text{Si}_3\text{N}_4$ ) is made by combining Si and nitrogen gas ( $\text{N}_2$ ) at a high temperature. How much (in g) Si is needed to react with an excess of nitrogen gas to prepare 125 g of silicon nitride if the percent yield of the reaction is 95.0%? 79.1 g Si  
Souring of wine occurs when ethanol is converted to acetic acid by oxygen by the following reaction:  
 $\text{C}_2\text{H}_5\text{OH} + \frac{1}{2}\text{O}_2 \rightarrow \text{CH}_3\text{COOH} + \text{H}_2\text{O}$   
A 1.00 L bottle of wine, labeled as 8.5% (by volume) ethanol, is found to have a defective seal. Analysis of 1.00 mL showed that there were 0.0274 grams of acetic acid in that 1.00 mL. The density of ethanol is 0.816 g/mL, and the density of water is 1.00 g/mL.  
a. What mass of oxygen must have leaked into the bottle? 14.6 g  
b. What is the percent yield for the conversion of ethanol to acetic acid if  $\text{O}_2$  is in excess? 30.3%  
A reaction container holds 5.77 g of  $\text{P}_4$  and 5.77 g of  $\text{O}_2$ . The following reaction occurs:  
 $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_4\text{O}_6$   
If enough oxygen is available then the  $\text{P}_4\text{O}_6$  reacts further:  
 $\text{P}_4\text{O}_6 + \text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$   
a. What is the limiting reagent for the formation of  $\text{P}_4\text{O}_{10}$ ?  $\text{O}_2$   
b. What mass of  $\text{P}_4\text{O}_{10}$  is produced? 5.78 g  
c. What mass of excess reactant is left in the reaction container? 5.76 g  
 $\text{P}_4\text{O}_6$   
remain  
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